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| **Karan Arora** **R.L. Chemistry Classes M: 99968-68554**  **Class : X**  **“Chemical Reactions and Equations”** |

**Worksheet – 1**

1. Which of the following observation/s can be used to determine whether a chemical reaction has taken place ?

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| a) Change in colour | b) Change in temperature |
| c) Evolution of a gas | d) any one of the three |

1. Which of the following does not represent the balanced chemical reaction correctly ?
2. Na (s) H2O (l) → NaOH (aq) + H2 (g).
3. BaCl2 (aq) + Na2SO4 (aq) → BaSO4 (aq) + 2 NaCl (aq).
4. H2 (g) + Cl2 (g) → 2 HCl (g).
5. CaO (s) + H2O (l) → Ca(OH)2 (aq)
6. The reaction in which heat is evolved are called

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| a) thermal reactions | b) exothermic reactions |
| c) endothermic reactions | d) photochemical reactions |

1. What is balanced chemical equation? Why should the chemical equation be balanced ?
2. What do you mean by exothermic and endothermic reactions? Give examples.
3. Why is respiration considered as an exothermic reaction? Explain.
4. Balance the following chemical equations :
5. Hydrogen + chlorine → Hydrogen chloride
6. Barium chloride + Aluminium sulphate → Barium sulphate + Aluminium chloride
7. Balance the following chemical equations :
8. Al + HCl → AlCl3 + H2
9. Mg + CO2 MgO + C
10. H2S + O2 → SO2 + H2O
11. NH3 + O2 → N2 + H2O
12. BaCl2 + H2SO4 → BaSO4 + HCl
13. NaOH + H2SO4 → Na2SO4 + H2O